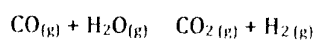


20. The following table gives some values for reactant and product equilibrium concentrations (in moles/L; M) at 700 K for the Shift reaction, an important method for the commercial production of hydrogen gas:



Trial	[CO ₂]	[H ₂]	[CO]	[H ₂ O]
1	0.600	0.600	0.266	0.266
2	0.600	0.800	0.330	0.286
3	2	2	0.877	0.877
4	1.00	1.50	0.450	0.655
5	1.80	2.00	0.590	1.20

Calculate K_{eq} for each of the five trials. How do the answers compare with each other? Why?

$$K = 5.09$$